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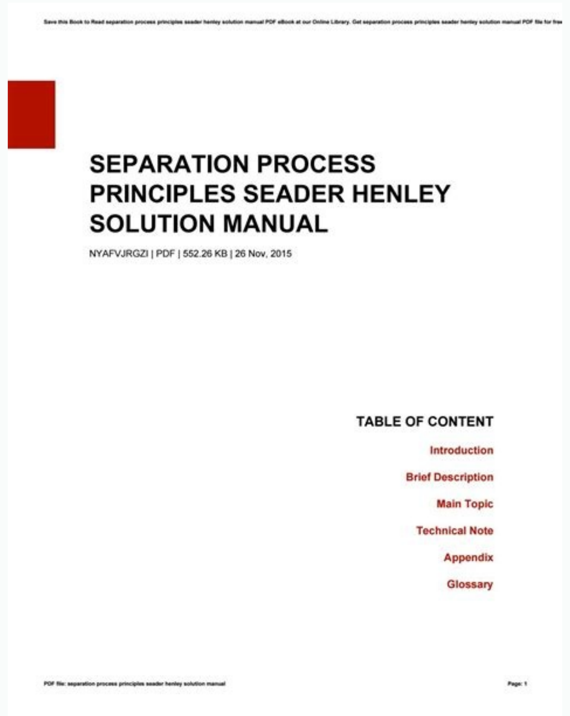
Separation process principles solutions manual answers pdf online test



SEPARATION PROCESS PRINCIPLES SEADER HENLEY SOLUTION MANUAL

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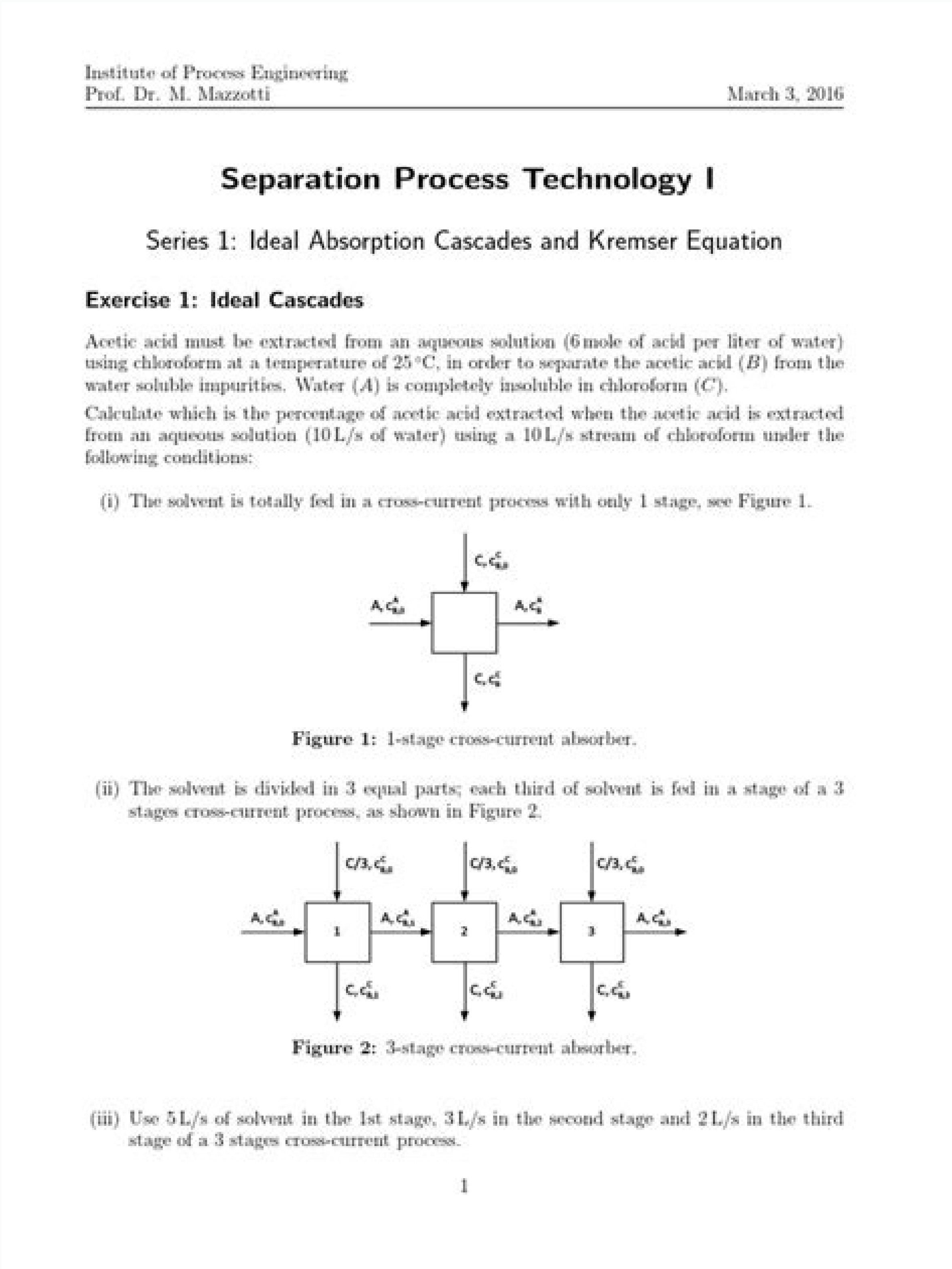


Separation Process Technology I

Series 1: Ideal Absorption Cascades and Kremser Equation

Exercise 1: Ideal Cascades

Acetic acid must be extracted from an aqueous solution (6 mole of acid per liter of water) using chloroform at a temperature of 25 °C, in order to separate the acetic acid (B) from the water soluble impurities. Water (A) is completely insoluble in chloroform (C).

Calculate which is the percentage of acetic acid extracted when the acetic acid is extracted from an aqueous solution (10 L/s of water) using a 10 L/s stream of chloroform under the following conditions:

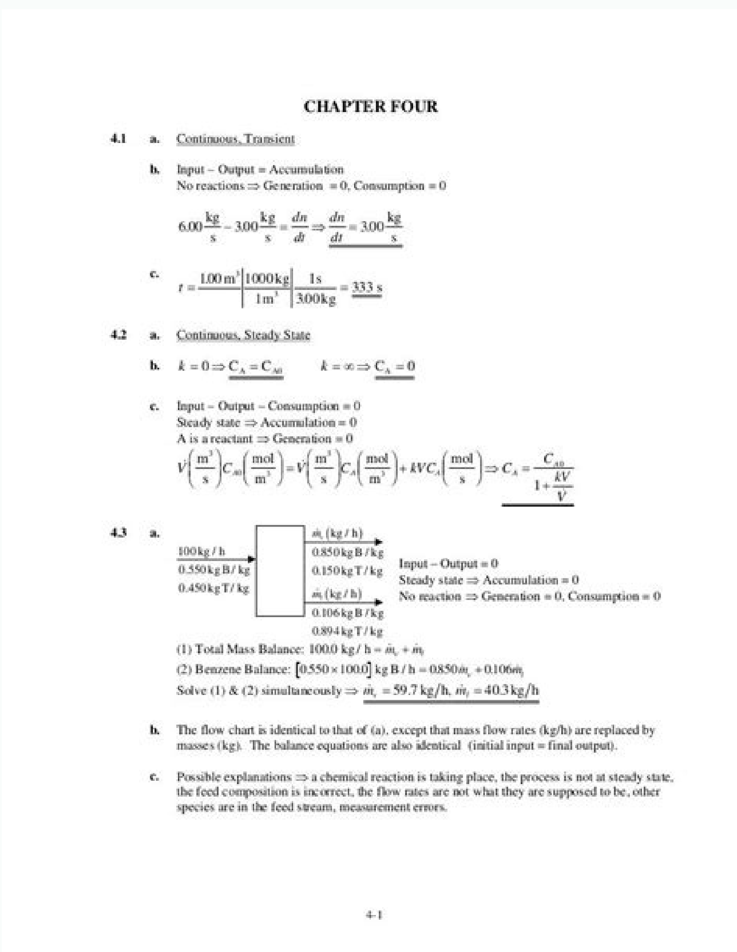
(i) The solvent is totally fed in a cross-current process with only 1 stage, see Figure 1.

Figure 1: 1-stage cross-current absorber.

(ii) The solvent is divided in 3 equal parts; each third of solvent is fed in a stage of a 3 stages cross-current process, as shown in Figure 2.

Figure 2: 3-stage cross-current absorber.

(iii) Use 5 L/s of solvent in the 1st stage, 3 L/s in the second stage and 2 L/s in the third stage of a 3 stages cross-current process.



CHAPTER FOUR

4.1 a. Continuous, Transient

b. Input − Output = Accumulation
No reactions ⇒ Generation = 0, Consumption = 0

$$6.00\frac{kg}{s} - 3.00\frac{kg}{s} = \frac{dn}{dt} \Rightarrow \frac{dn}{dt} = 3.00\frac{kg}{s}$$

c. $t = \frac{1.00\,m^3}{} \Big| \frac{1000\,kg}{1\,m^3} \Big| \frac{1\,s}{3.00\,kg} = 333\,s$

4.2 a. Continuous, Steady State

b. $k = 0 \Rightarrow C_A = C_{A0}$ $k = \infty \Rightarrow C_A = 0$

c. Input − Output − Consumption = 0
Steady state ⇒ Accumulation = 0
A is a reactant ⇒ Generation = 0

$$\dot{V}\left(\frac{m^3}{s}\right)C_{A0}\left(\frac{mol}{m^3}\right) = \dot{V}\left(\frac{m^3}{s}\right)C_A\left(\frac{mol}{m^3}\right) + kVC_A\left(\frac{mol}{s}\right) \Rightarrow C_A = \frac{C_{A0}}{1+\frac{kV}{\dot{V}}}$$

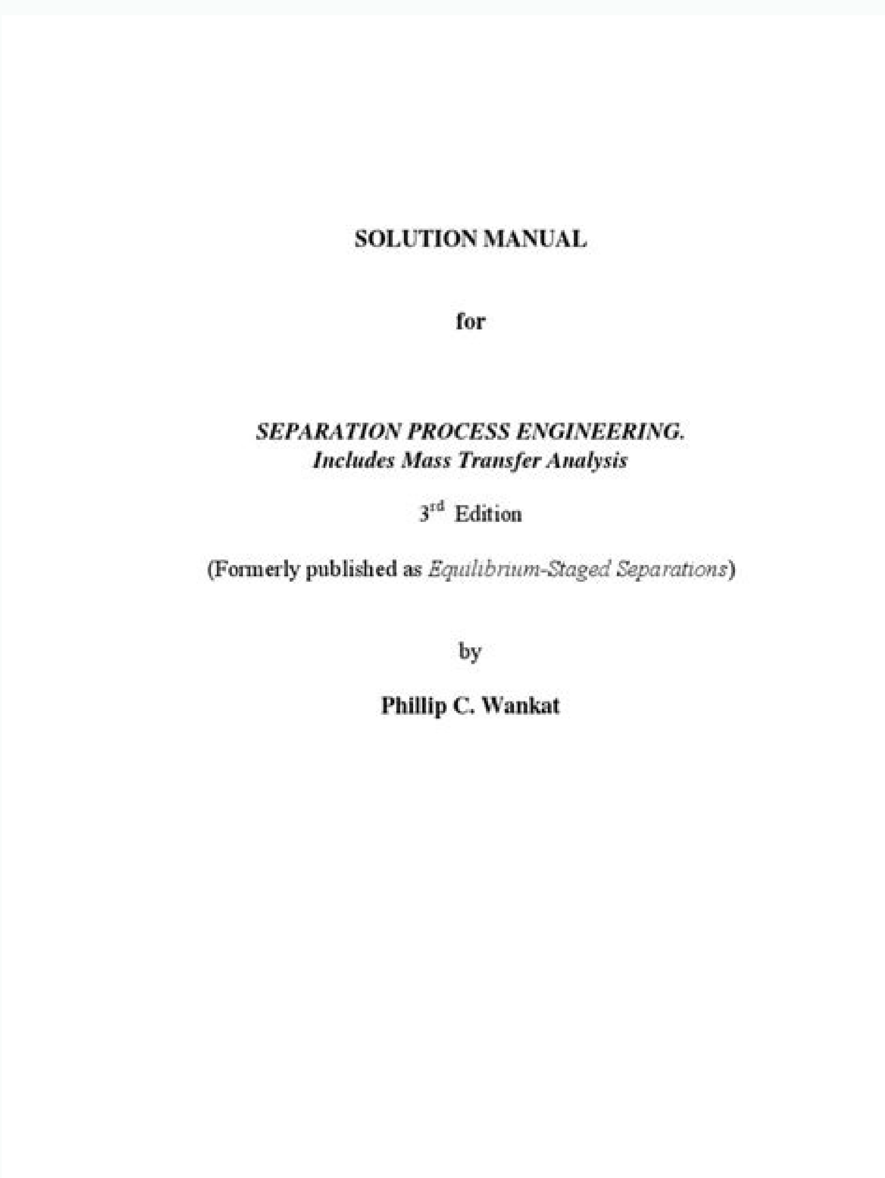
4.3 a. 100 kg/h, 0.550 kg B/kg, 0.450 kg T/kg → $\dot{m}_v$ (kg/h), 0.850 kg B/kg, 0.150 kg T/kg; $\dot{m}_l$ (kg/h), 0.106 kg B/kg, 0.894 kg T/kg

Input − Output = 0
Steady state ⇒ Accumulation = 0
No reaction ⇒ Generation = 0, Consumption = 0

(1) Total Mass Balance: $100.0\,kg/h = \dot{m}_v + \dot{m}_l$
(2) Benzene Balance: $[0.550 \times 100.0]\,kg\,B/h = 0.850\dot{m}_v + 0.106\dot{m}_l$
Solve (1) & (2) simultaneously ⇒ $\dot{m}_v = 59.7\,kg/h$, $\dot{m}_l = 40.3\,kg/h$

b. The flow chart is identical to that of (a), except that mass flow rates (kg/h) are replaced by masses (kg). The balance equations are also identical (initial input = final output).

c. Possible explanations ⇒ a chemical reaction is taking place, the process is not at steady state, the feed composition is incorrect, the flow rates are not what they are supposed to be, other species are in the feed stream, measurement errors.



SOLUTION MANUAL

for

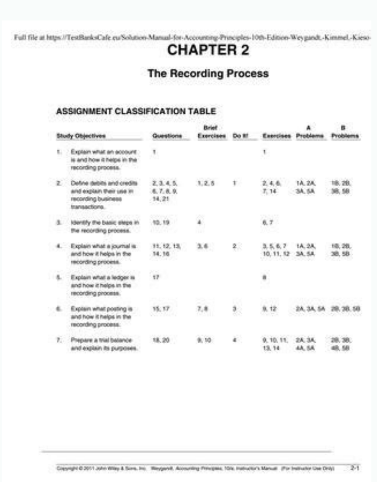
SEPARATION PROCESS ENGINEERING.
Includes Mass Transfer Analysis

3rd Edition

(Formerly published as *Equilibrium-Staged Separations*)

by

Phillip C. Wankat



CHAPTER 2
The Recording Process

ASSIGNMENT CLASSIFICATION TABLE

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